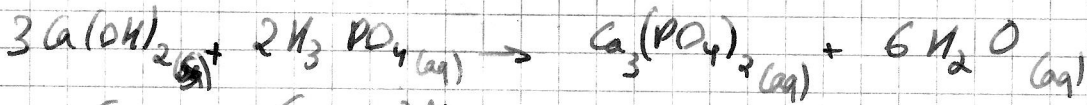


Reidetur in proba studio

1



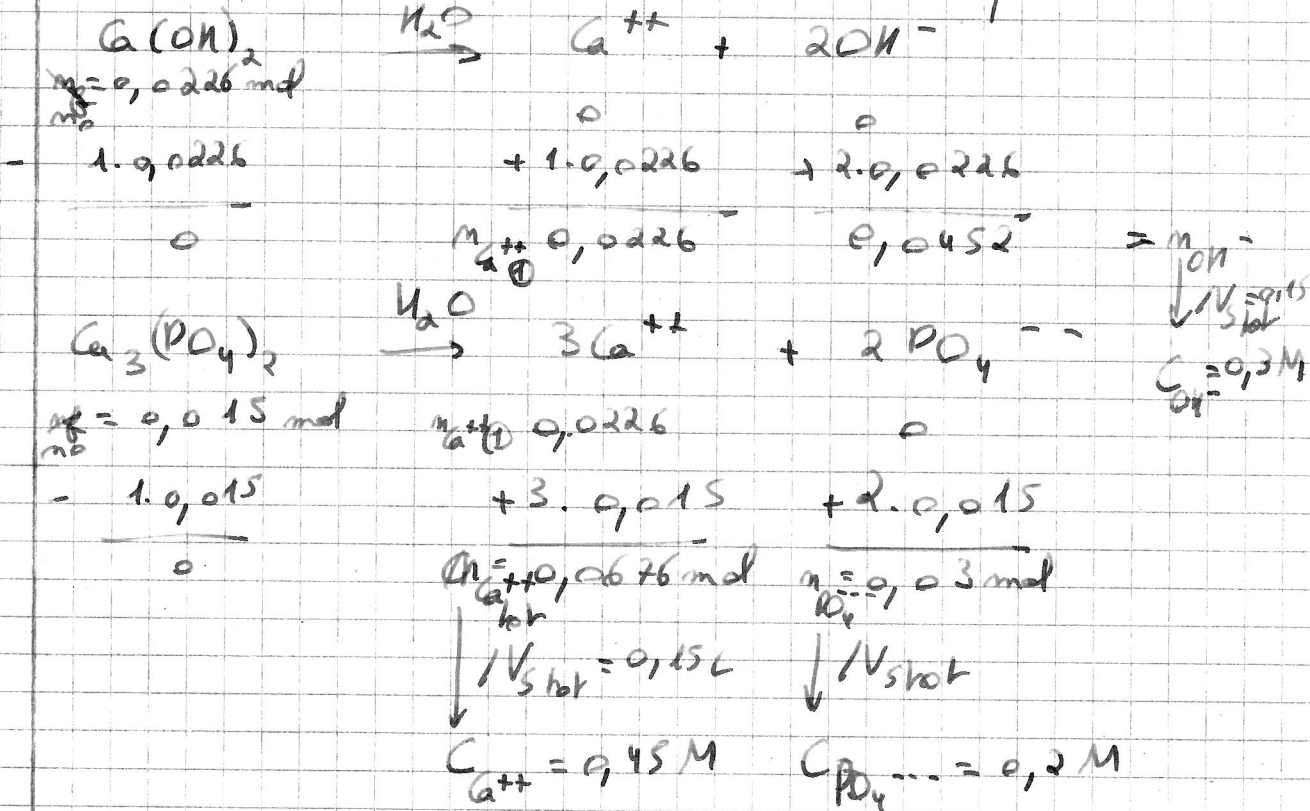
$m = 5 \text{ g}$   
 $\downarrow M = 74 \text{ g/mol}$   
 $n = 0,0676$

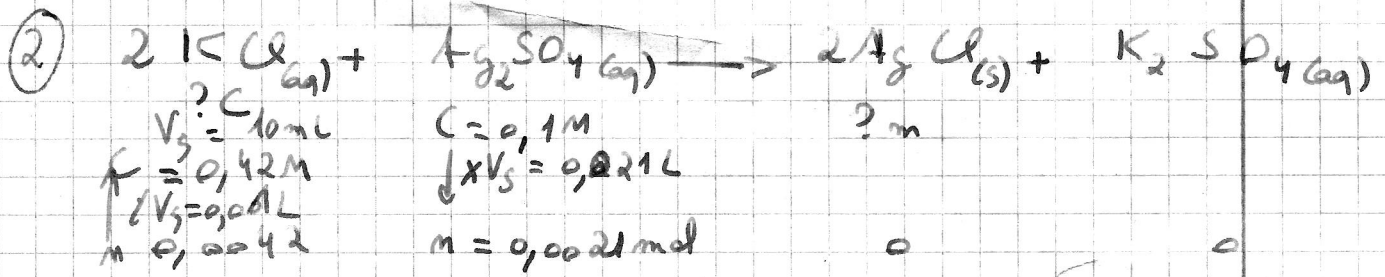
$C = 0,2 \text{ M}$   
 $\downarrow \times V_s = 150 \text{ mL}$   
 $m = 0,03 \text{ mol}$

$\Delta n$	-3.0,015	-2.0,015	+1.0,015	+6.0,015
$n_f$	0,0226	0	0,015	0,09

$\downarrow M = 310 \text{ g/mol}$   
 $m = 4,65 \text{ g}$   
 $\downarrow V_s = 0,15 \text{ L}$   
 $\rho = 31 \text{ g/L}$

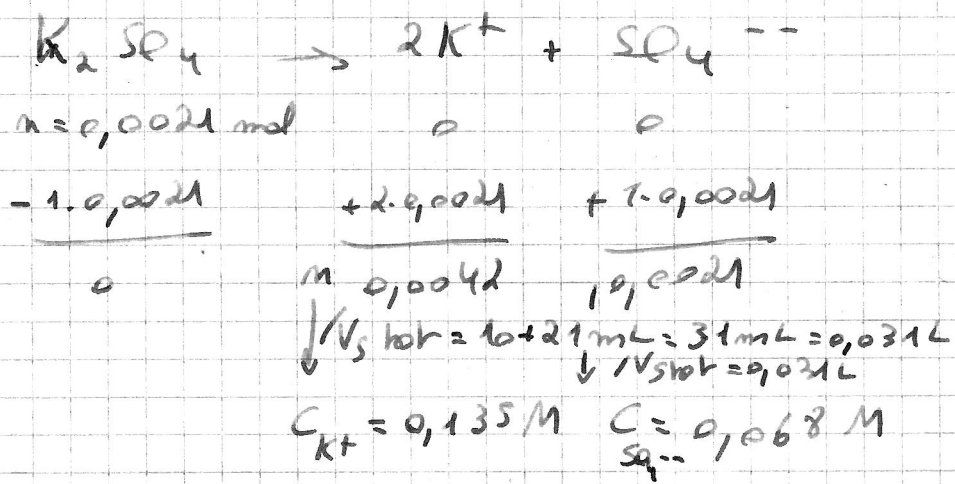
Quelle conc en tous les ions dans la solution finale?





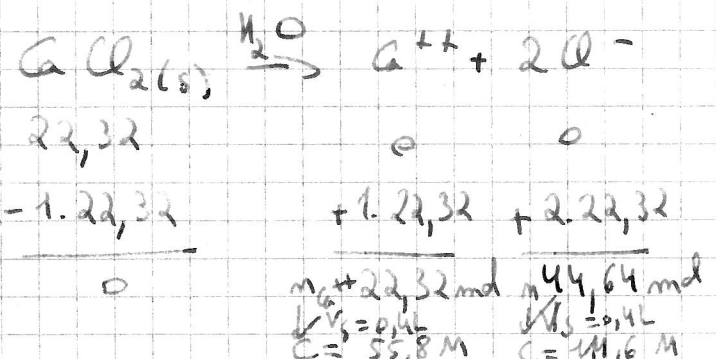
$-2 \cdot 0,002$	$-1 \cdot 0,002$	$+2 \cdot 0,002$	$+1 \cdot 0,002$
<hr/>	<hr/>	<hr/>	<hr/>
0	0	$n = 0,004 \text{ mol}$	$0,002$
Litrage = juste ce qu'il faut pour une réaction complète		$\downarrow \times M = 143,5 \text{ g/mol}$ $m = 0,603 \text{ g}$	

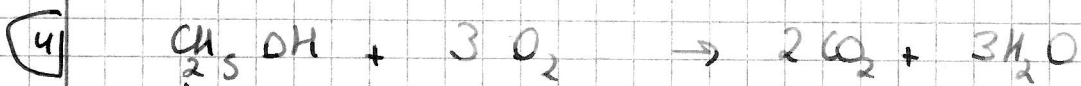
Quelles sont les C de tous les ions dans la solution finale



$V_{\text{gaz}} = 1 \text{ m}^3$   
 $= 1000 \text{ L}$   
 $\downarrow V_m = 22,4 \text{ L/mol}$   
 $n = 44,64 \text{ mol}$

$m = 1651,7 \text{ g}$			
$\uparrow M = 74 \text{ g/mol}$			
$n_0 = 22,32 \text{ mol}$			
$\downarrow n = 1 \cdot 22,32$		$-2 \cdot 22,32$	$+1 \cdot 22,32$
<hr/>		<hr/>	<hr/>
0		44,64	0
		$n = 22,32 \text{ mol}$	$n = 22,32$
		$\downarrow \times M = 18 \text{ g/mol}$	<del><math>\downarrow \times M</math></del>
		$m = 401,76 \text{ g}$	
		$\downarrow \rho_{\text{eau}} = 1 \text{ g/mL}$	
		$V_{\text{air}} = 401,76 \text{ mL}$	





Combustion  
= réaction avec  
 $\text{O}_2$  et on  
obtient les  
atomes des  
combustibles  
oxydés  
C O  
II II  
H<sub>2</sub>O  
I 2 II  
simplifier  
avec  $\text{O}_2$   
des  
réactifs

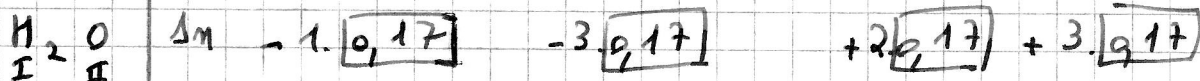
$V_{\text{liq}} = 10 \text{ mL}$

$\rho = 0,789 \text{ g/mL}$

$m = 7,89 \text{ g}$

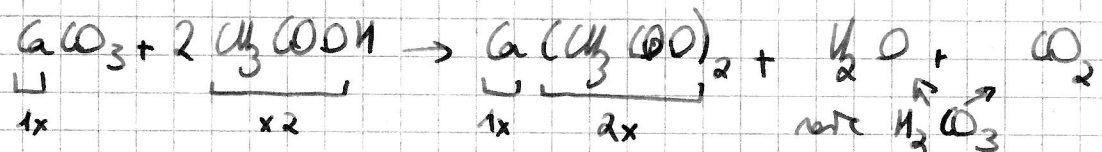
$M_{\text{CH}_3\text{OH}} = 46 \text{ g/mol}$

$n = 0,17 \text{ mol}$



$V_{\text{m}} = 22,4 \text{ L/mol}$   
 $V_{\text{gaz}} = 7,616 \text{ L}$

5)



Vinaigre 8% :  $m = 8 \text{ g}$  dans  $m_s = 100 \text{ g}$

$M_{\text{CH}_3\text{COOH}} = 60 \text{ g/mol}$

$m = 0,1333 \text{ mol}$

$V_s = 0,1 \text{ L}$

$C = 1,333 \text{ M}$

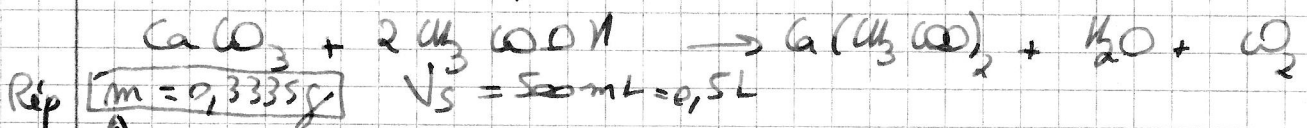
ici considéré  
comme de l'eau  
car annoncé 80 g/L

$\rho_{\text{eau}} = 1 \text{ g/mL}$   
 $V_s = 100 \text{ mL}$

autre info :  $\rho = 80 \text{ g/L}$

$M_{\text{CH}_3\text{COOH}} = 60 \text{ g/mol}$

$C = 0,1333 \text{ M}$



$xM = 0,003335 \text{ mol}$

$C = 0,1333 \text{ M}$

$xV_s = 0,5 \text{ L}$

$n = 0,003335$

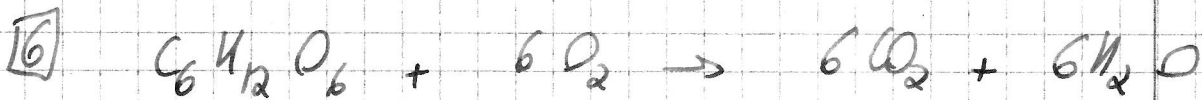
$m = 0,0667 \text{ mol}$

$\Delta n = -1 \cdot 0,003335$

$-2 \cdot 0,003335$

$n_f = 0$

$0$



$$m = 2,2 \text{ g}$$

$$\downarrow / M_{\text{CO}_2} = 44 \text{ g/mol}$$

$$n = 0,05 \text{ mol}$$

$n_0$	0,0083	0,05	0	0
$\Delta n$	-1. $\boxed{0,0083}$	-6. $\boxed{0,0083}$	+6. $\boxed{0,0083}$	+6. $\boxed{0,0083}$
$n_f$	0	0	0,05	0,05

$$n_0 = 0,0083 \text{ mol} \quad n_0 = 0,05 \text{ mol}$$

$$\downarrow \times M_{\text{C}_6\text{H}_{12}\text{O}_6} = 180 \text{ g/mol} \quad \downarrow \times M_{\text{O}_2} = 32 \text{ g/mol}$$

$$m = 1,5 \text{ g}$$

$$m = 1,6 \text{ g}$$

$$n_f = 0,05 \text{ mol}$$

$$\downarrow \times M_{\text{H}_2\text{O}} = 18 \text{ g/mol}$$

$$m = 0,9 \text{ g}$$

3,1 g de réactifs

3,1 g de produits