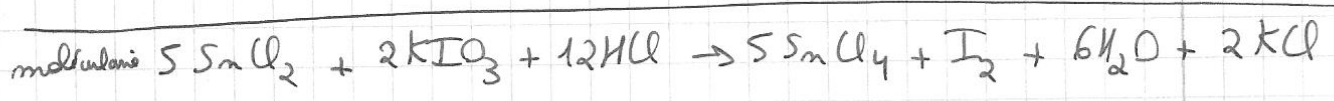
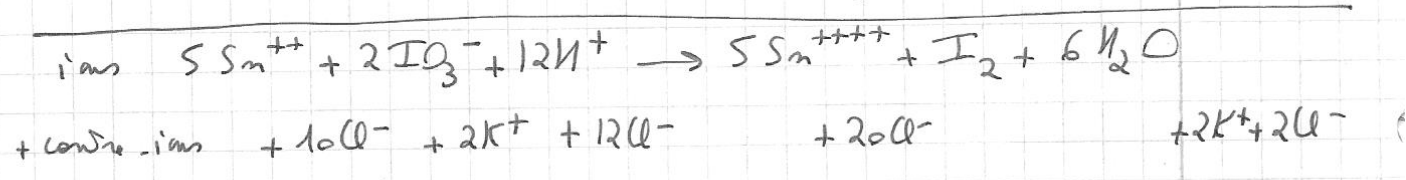
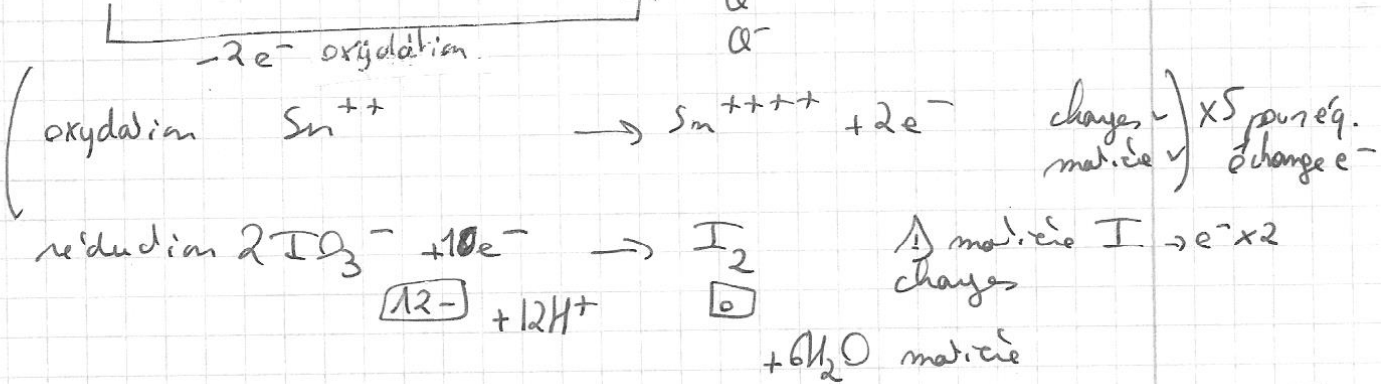
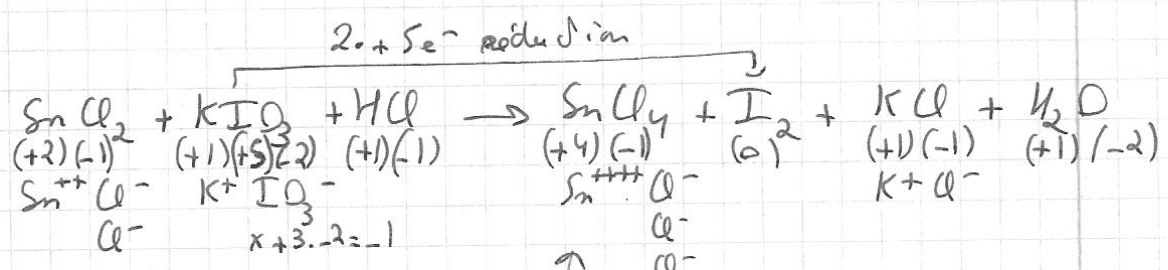


M



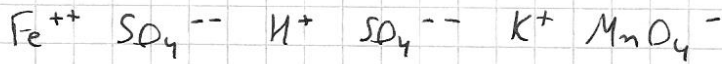
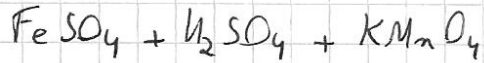
V_S = 20 mL V_S = 30 mL

C = 0,01 M
 ↓ x V_S = 0,03 L

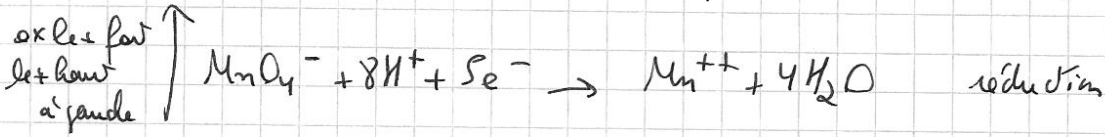
m ₀	0,00075	m = 0,0003 mol
Δm	-5,00015	-2,00015
mf	0	0

m = 0,00075 mol
 ↓ V_S = 0,02 L
 C = 0,0375 M

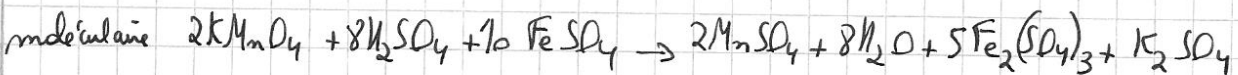
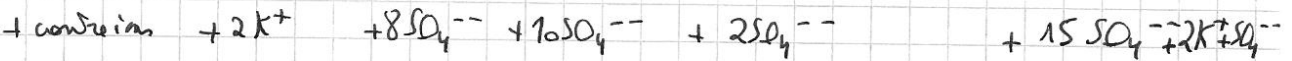
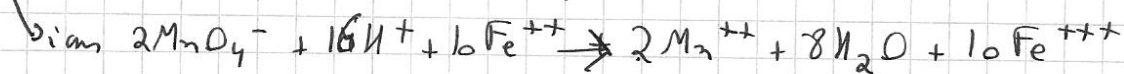
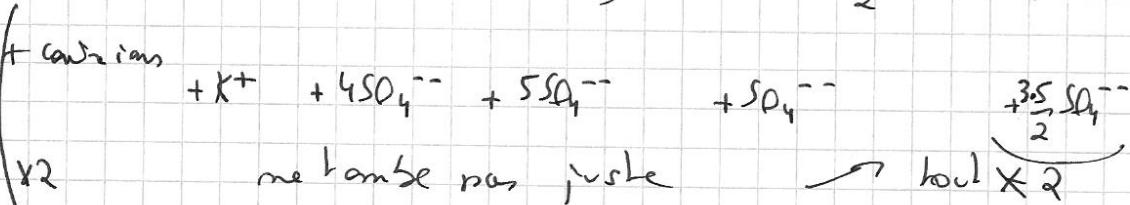
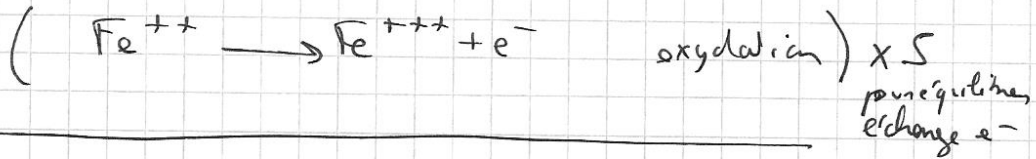
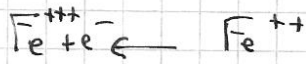
12



recherche dans tableau de oxydant le + fort et reducant le + fort



red le + fort
le + bas
à droite



$$V_S = 52 \text{ mL}$$

$$m = 19,7 \text{ g}$$

$$\downarrow V_S = 1 \text{ L}$$

$$\rho = 19,7 \text{ g/L}$$

$$\downarrow \times V_S = 10 \text{ mL} = 0,01 \text{ L}$$

$$m = 0,197 \text{ g}$$

$$\downarrow M_{\text{FeSO}_4} = 56 + 32 + 4 \cdot 16 = 152 \text{ g/mol}$$

$$n = 0,0013 \text{ mol}$$

$$n_0 = 0,00026$$

$$\Delta n = -2 \cdot 0,00013$$

$$-10 \cdot 0,00013$$

∑

0

0

$$n = 0,00026 \text{ mol}$$

$$\downarrow V_S = 0,052 \text{ L}$$

$$C = 0,005 \text{ M}$$

