

Chap 12

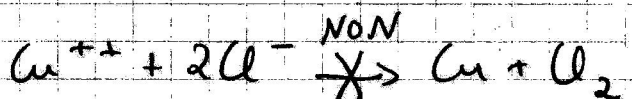
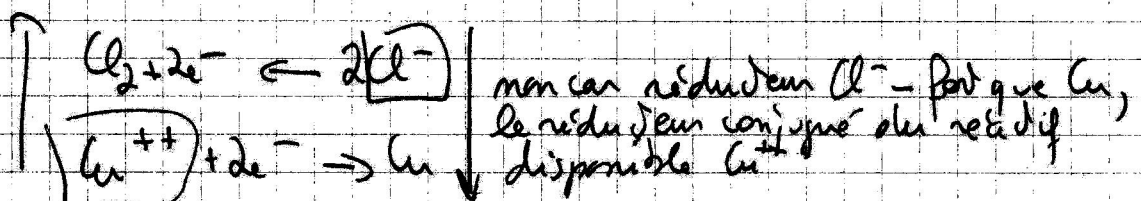
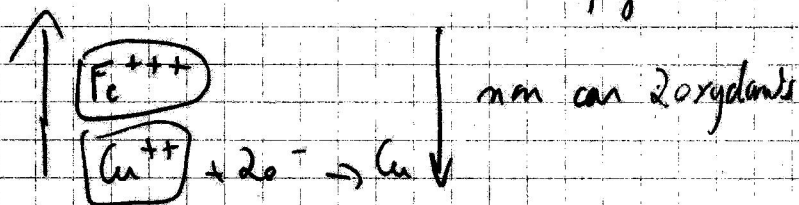
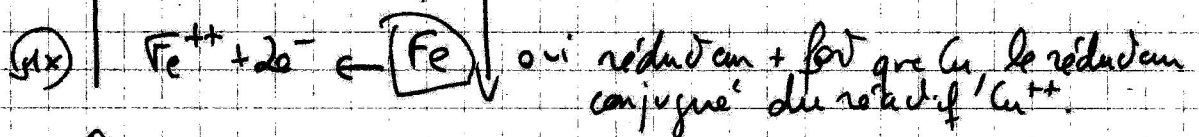
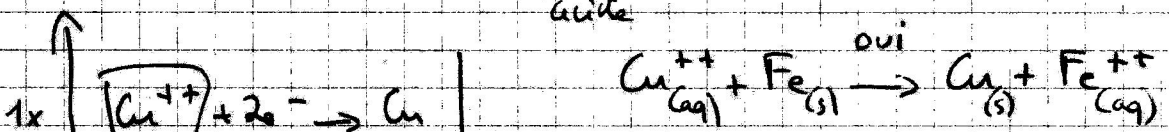
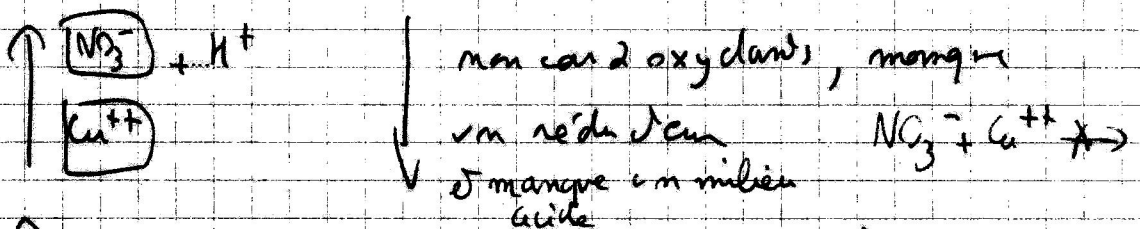
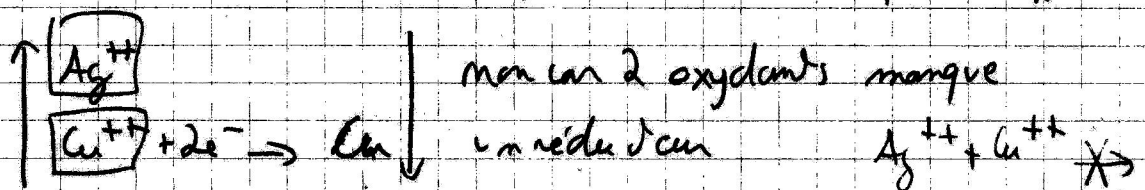
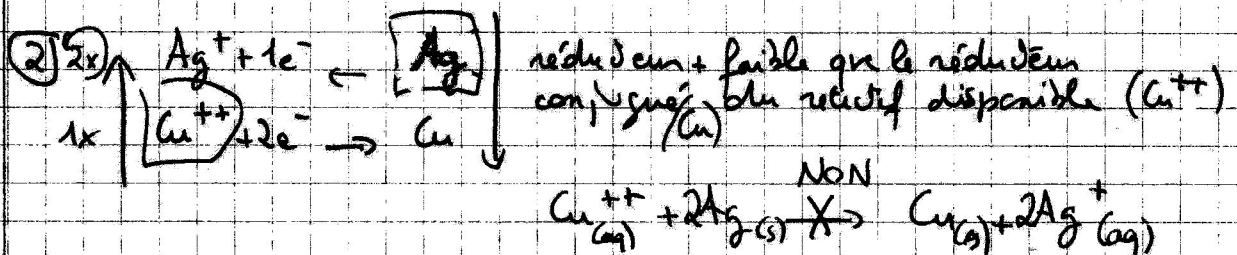
① Cl^- : très réducteur

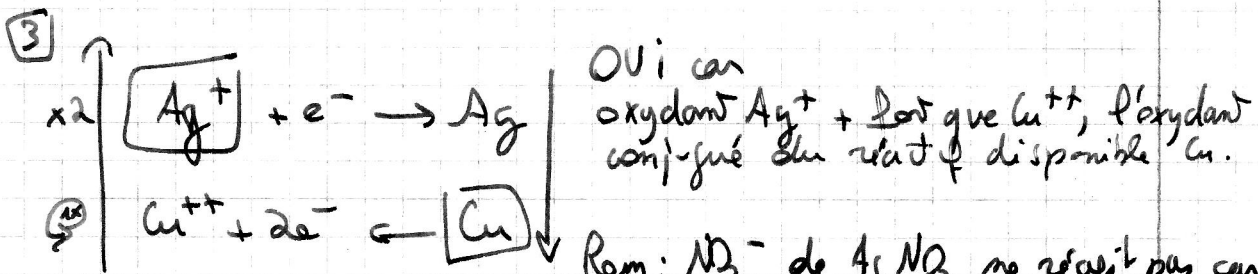
car il réapparaît que à droite dans la équation de potentiel de réduction du tableau.

ou car dans la colonne des halogènes : VIIa ~~XI~~

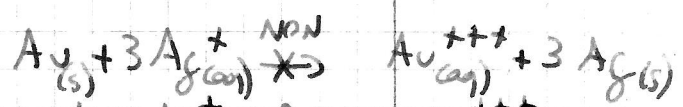
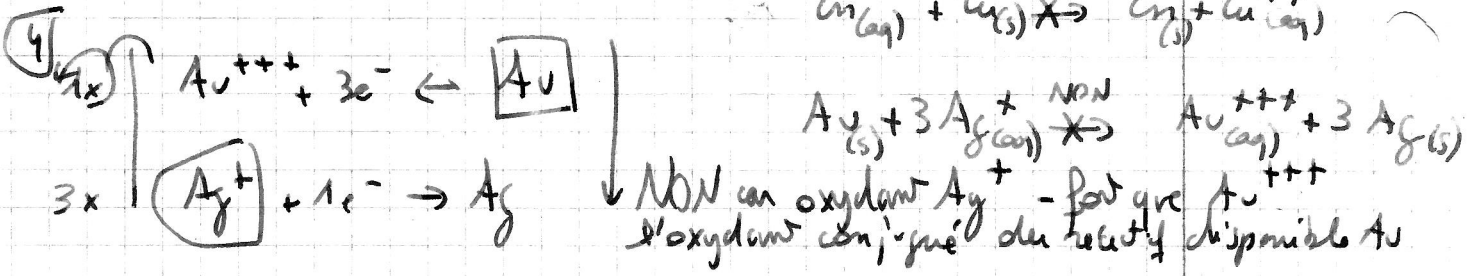
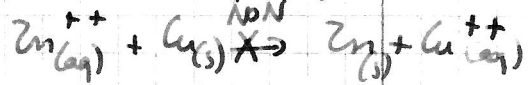
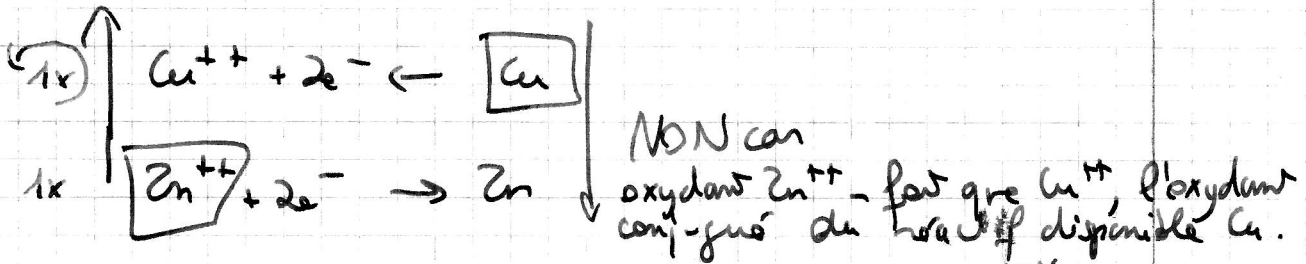
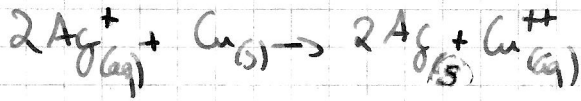
les atomes peuvent avoir un NO entre -1 : couche complétée et +7 : couche vidée

Cl^- (C-1) ne peut plus gagner d' e^- ; il ne peut qu'en perdre : subir une oxydation et c'est donc toujours un réducteur.

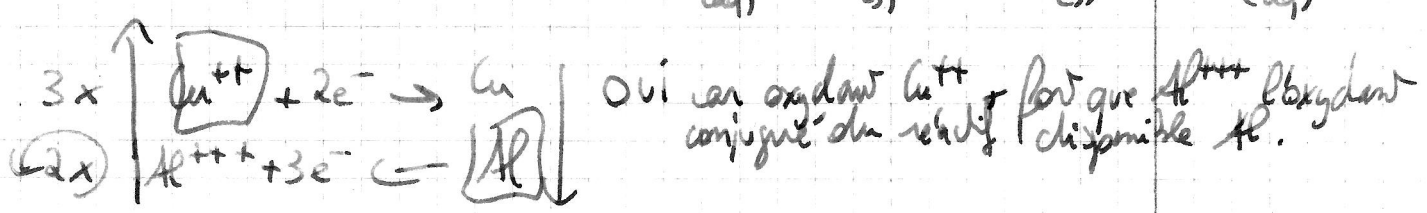
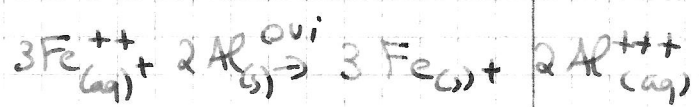
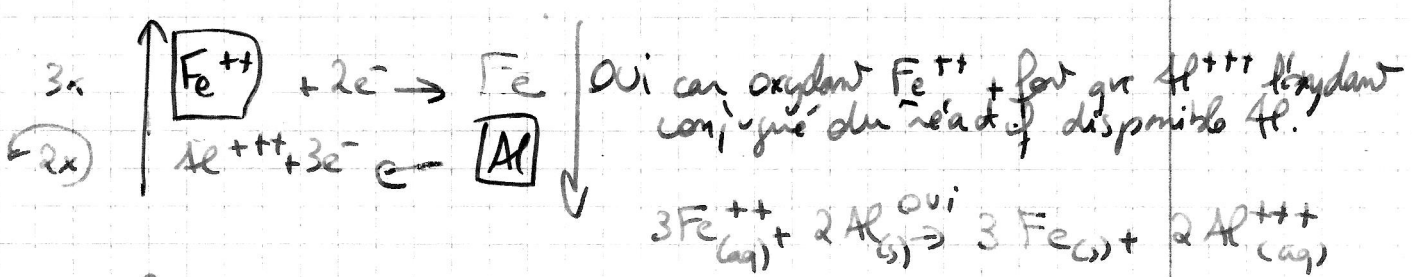
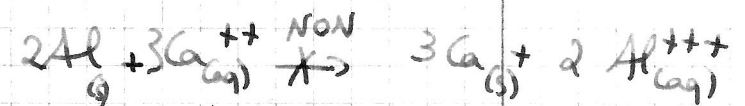
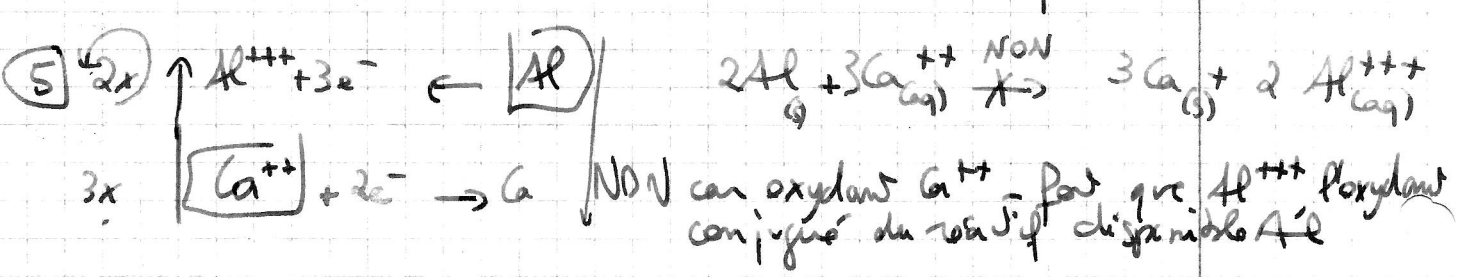


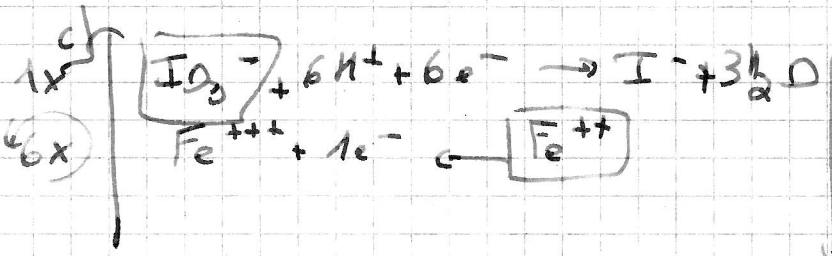
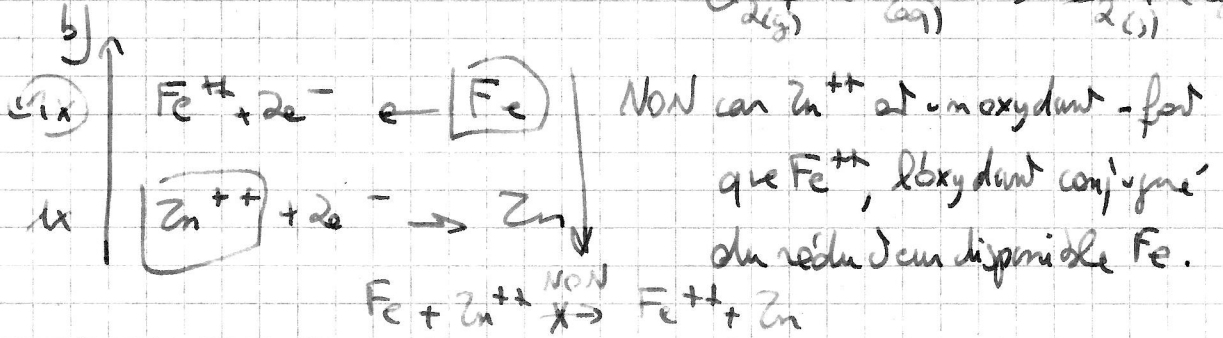
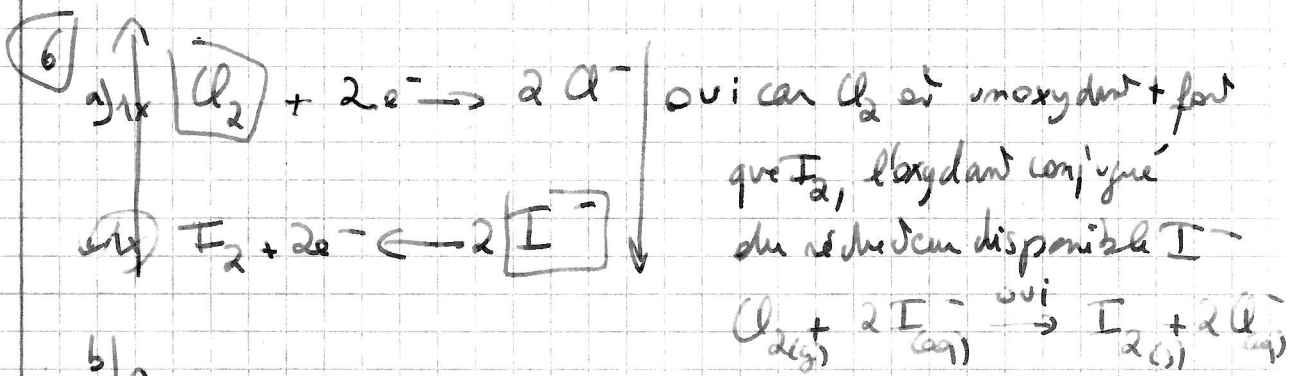


Rem: NO_3^- de $AgNO_3$ ne réagit pas car il n'y a pas d'acide dans la solution $\rightarrow NO_3^- + \dots H^+ \rightarrow \dots$ besoin acide.



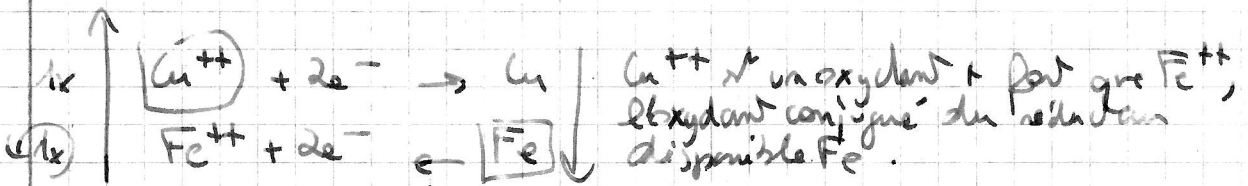
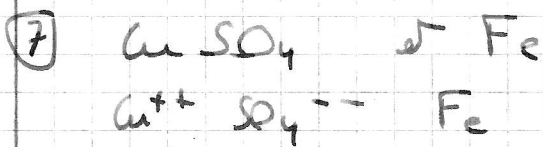
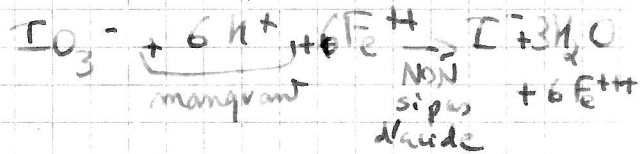
Rem: NO_3^- est aussi un oxydant - fort que Au^{+++} et il demande un milieu acide non fourni ici.



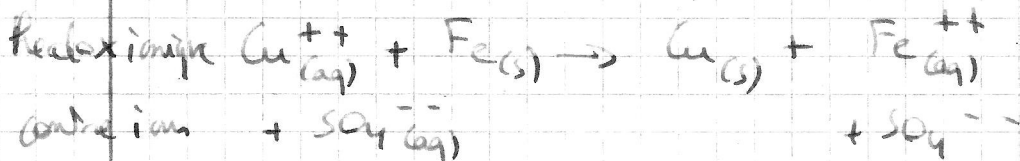


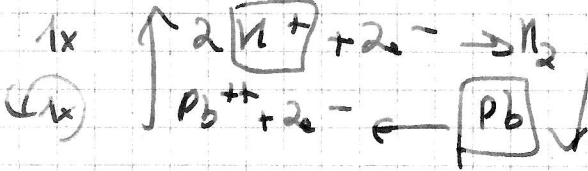
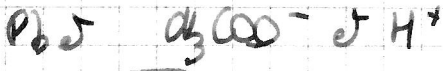
La force de l'oxydant et du réducteur le permettrait
 respectivement

mais IO_3^- nécessite un milieu acide qui n'est pas
 précisé ici donc NON.



Le Cu^{++} de la solution formera du Cu
 et le Fe qui réagira passera en solution, on inverse les lieux.



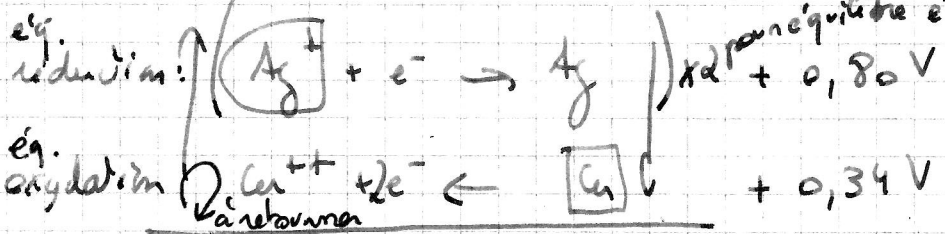
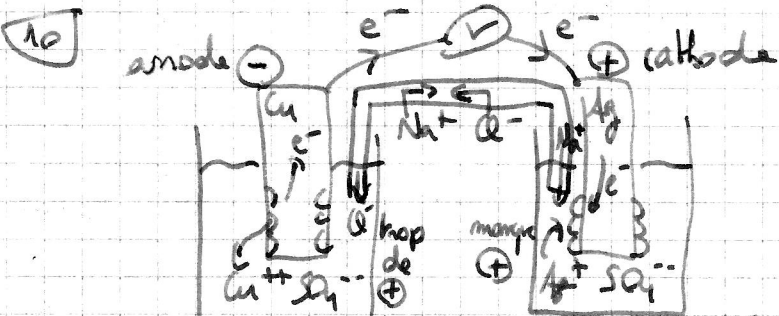
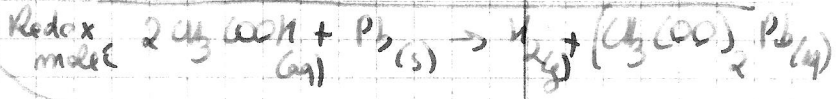
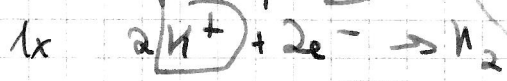


H⁺ est un oxydant + fort que Pb⁺⁺
 l'oxydant conjugué du réducteur disponible Pb.

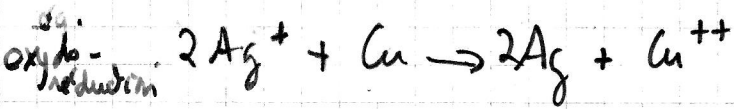
L'H⁺ retirera une couche de Pb des soldes: attaque.

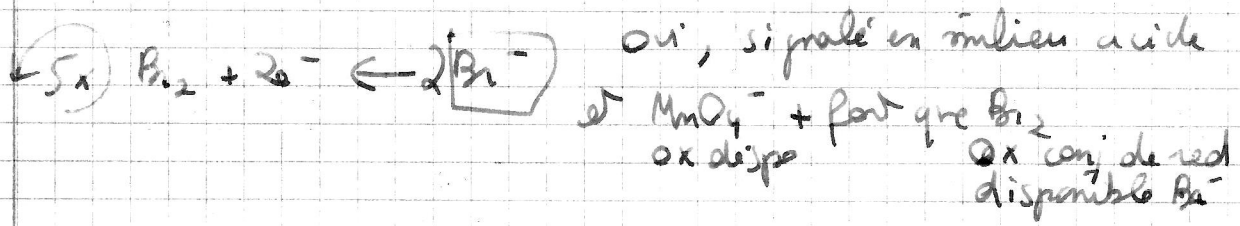
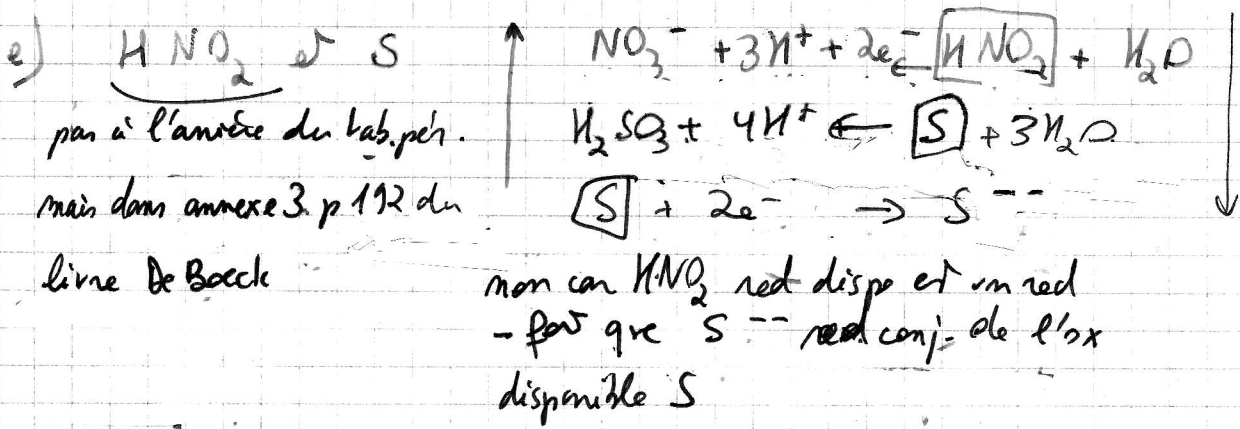
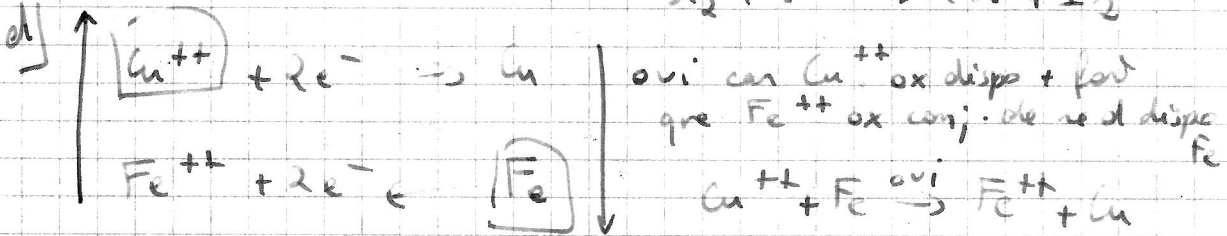
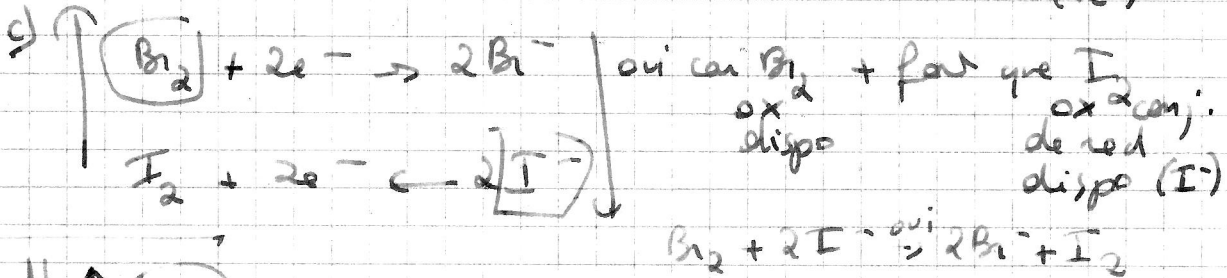
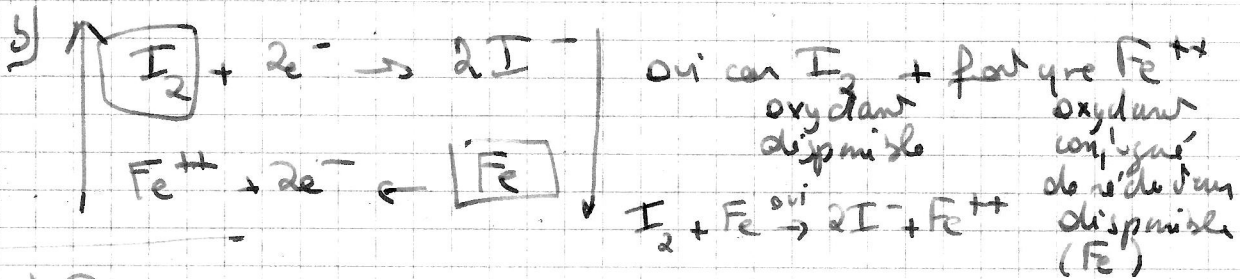
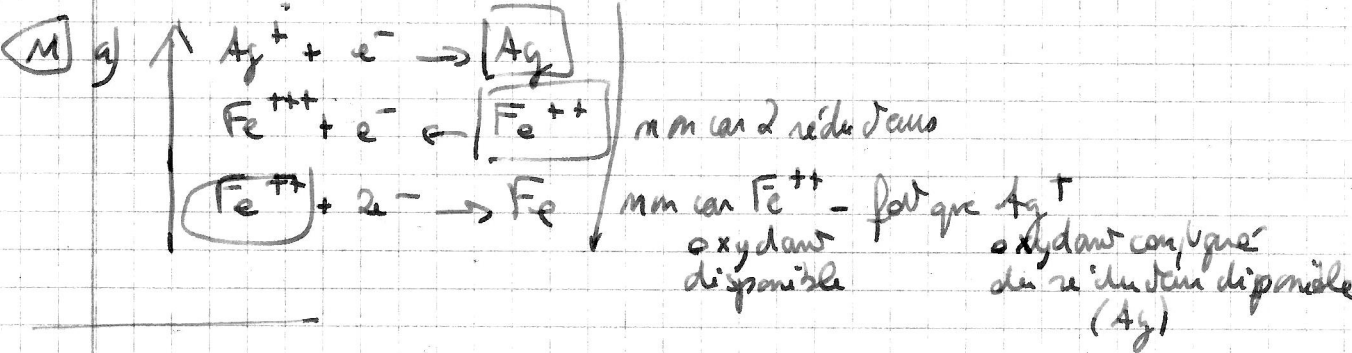


9 m couples que 8

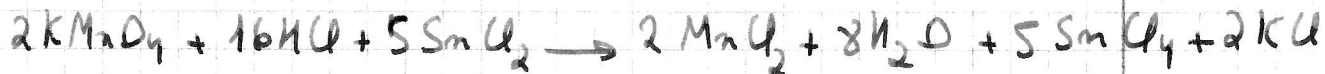
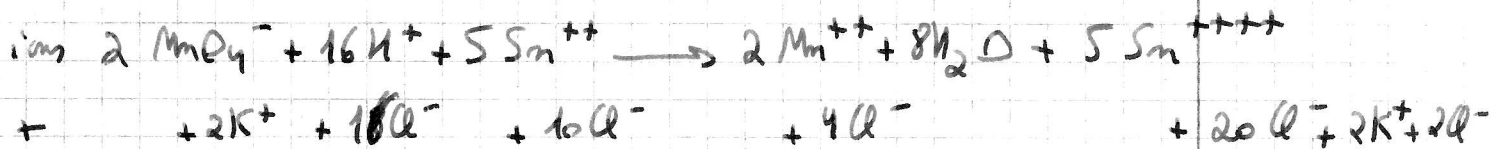
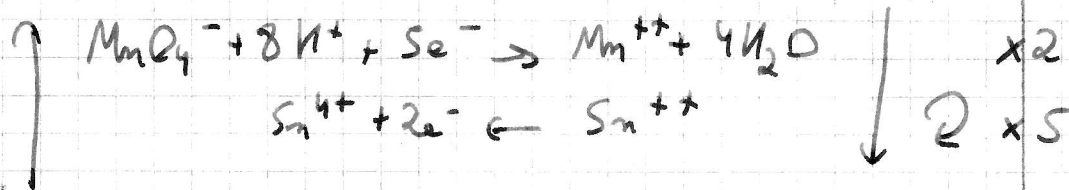


$\Delta E_0^0 = E_1^0 - E_2^0$
 $= 0,80 - 0,34$
 $= 0,46 V$





12) SnCl_2 KMnO_4 , milieci acide



$$m = 72 \text{ g}$$

$$C = 1 \text{ M}$$

$$M = 39 + 55 + 4 \cdot 16 = 158 \text{ g/mol}$$

$$m_{\text{Sn}} = 0,4557 \text{ mol}$$

$$1,13925$$

$$\Delta m = -2,0,22785$$

$$-5,0,22785$$

$$n_{\text{SnCl}_2} = 1,13925 \text{ mol}$$

$$? V_S \quad m = C \cdot V_S \quad \text{or} \quad C = \frac{m}{V_S}$$

$$V_S = \frac{m}{C} = \frac{1,13925}{1} = 1,13925 \text{ L}$$