

Chap 1

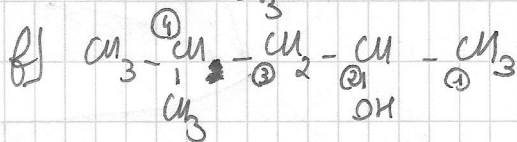
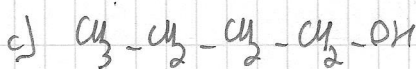
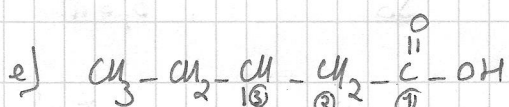
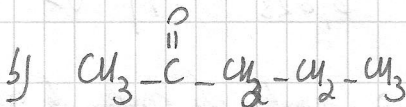
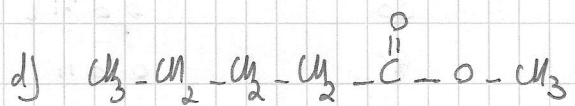
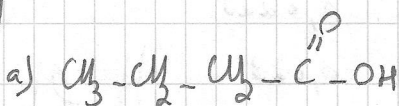
1. éter éthanoate de méthyle
 2. cétone
 3. alcool propan-1-ol
 4. éter propanoate d'éthyle
 5. acide acide butanoïque

c

1. combustion d'alcane : combustible + O₂ → atomes du combustible oxydés
 2. acide + alcool ⇌ éter + eau : estérification
 3. alcène + HCl → chloroalcane : addition
 4. éter + eau ⇌ ~~alco~~ol + alcool : hydrolyse
 5. alcène + H₂O → alcane : addition

d

3



4 a) acide propanoïque

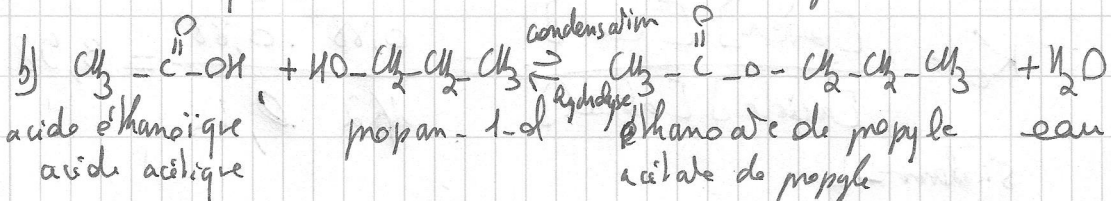
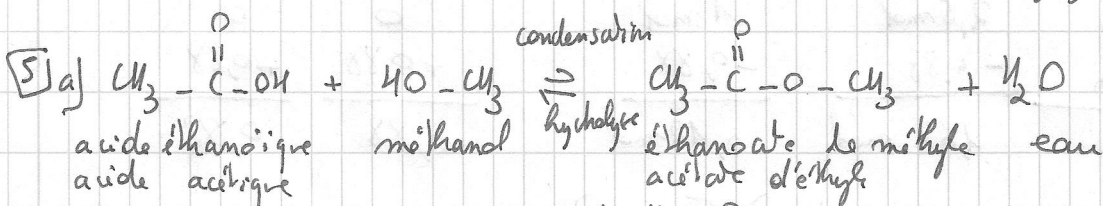
d) acide hexan-1,6-dioïque

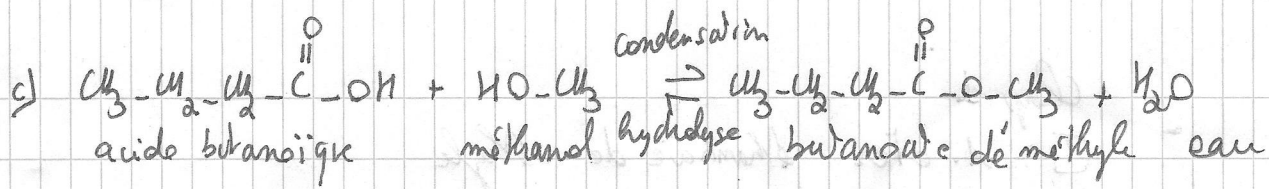
b) butanoate d'éthyle

e) éthanoate de propyle

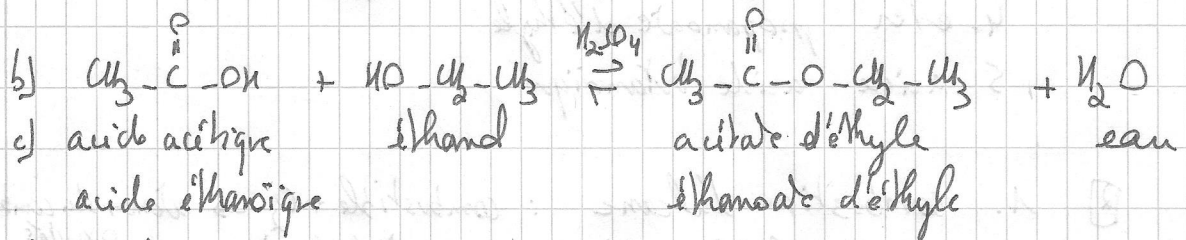
c) propan-1-ol

f) propan-1,2,3-triol
aussi appelé glycérol





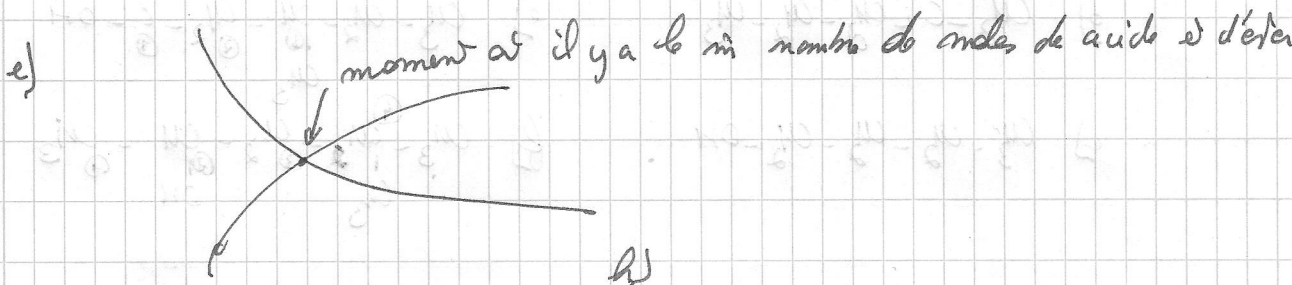
6) a) condensation, esterification



d)

t	naide (mol)	ester (mol)
0	0,034	0
5	0,0225	0,0115
10	0,018	0,016
15	0,013	0,021
30	0,015	0,0225
45	0,011	0,023
50	0,011	0,023

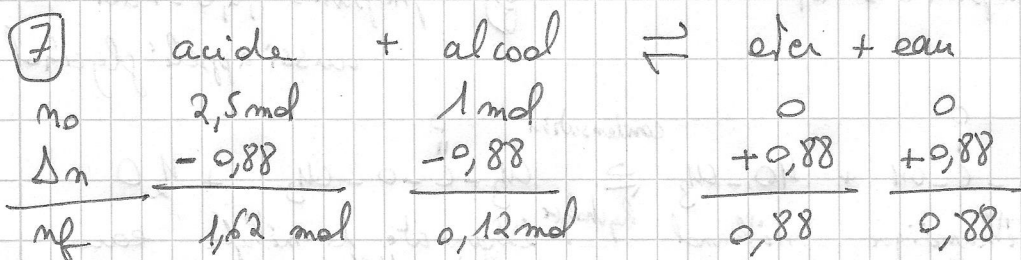
)) b)



f)

$$K_c = \frac{[\text{ester}][\text{eau}]}{[\text{acide}][\text{alcool}]} = \frac{0,023 \cdot 0,023}{0,031 \cdot 0,031} = 4,37 \quad \checkmark$$

si dans 1 l.



$$K_c = \frac{[\text{ester}][\text{eau}]}{[\text{acide}][\text{alcool}]} = \frac{0,88 \cdot 0,88}{1,62 \cdot 0,12} = 3,98 \quad \checkmark$$

si dans 1 L