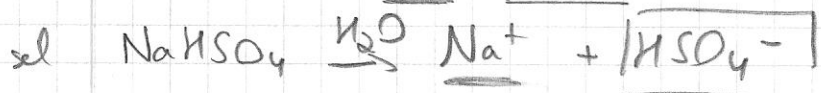
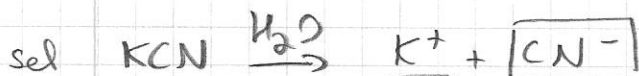
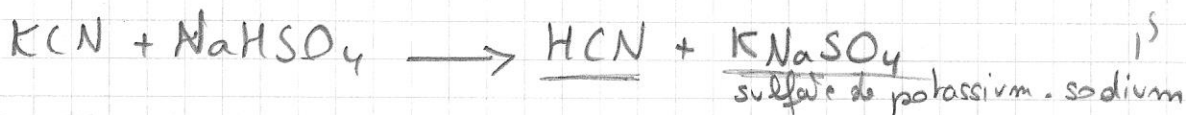


## KCN et NaHSO<sub>4</sub>

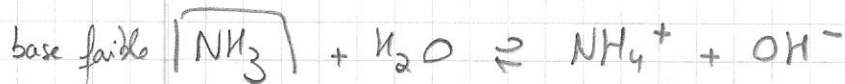
Concédant l'atome de carbone  
à l'oxygène



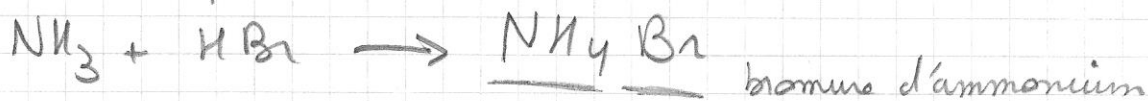
$$K_c = \frac{K_{aHSO_4^-}}{K_{aHCN}} = \frac{1,25 \cdot 10^{-2}}{5 \cdot 10^{-10}} = 2,5 \cdot 10^7 > 10^3 \text{ : complète}$$



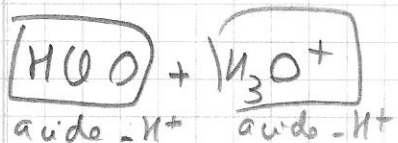
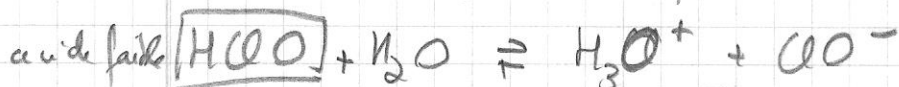
## NH<sub>3</sub> et HBr



$$K_c = \frac{K_{aH_3O^+}}{K_{aNH_4^+}} = \frac{55,5}{6 \cdot 10^{-10}} = 9,25 \cdot 10^{10} > 10^3 \text{ complète}$$



## HCOO et HCO<sub>2</sub>



besoin d'une base

pas de réaction entre 2 acides