

$$V_{\text{gas}} = 1 \text{ m}^3 = 1000 \text{ L}$$

$$N_m = 22,4 \text{ L/mol}$$

$$T = 20^\circ\text{C}$$

$$n_f = 44,6 \text{ mol}$$

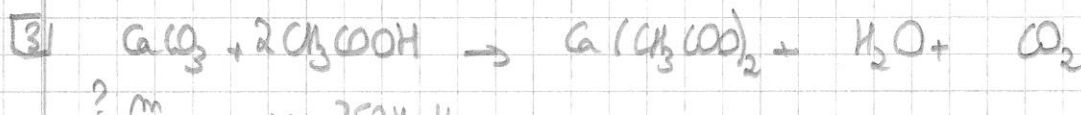
m_o	<u>22,3</u>	44,6	0	0	0
Δm	<u>-1,22,3</u>	<u>-2,22,3</u>	<u>+1,22,3</u>	<u>+1,22,3</u>	<u>+2,22,3</u>
n_f	0	0	22,3	22,3	<u>44,6</u>

$$m_o = 22,3 \text{ mol}$$

$$\downarrow \times M = 40 + 2 \cdot 16 + 2 \cdot 1$$

$$= 74 \text{ g/mol}$$

$$m = 1650,2 \text{ g}$$



$$\rho = 75,74 \text{ g/L}$$

$$\downarrow \times V_s = 500 \text{ mL} = 0,5 \text{ L}$$

$$m = 37,87 \text{ g}$$

$$\downarrow M = 2 \cdot 12 + 2 \cdot 16 + 4 \cdot 1$$

$$= 60 \text{ g/mol}$$

$$m_o = 0,63 \text{ mol}$$

m_o	<u>10,315</u>	<u>10,63</u>	0	0	0
Δm	<u>-1,0,315</u>	<u>-2,0,315</u>	<u>+1,0,315</u>	<u>+1,0,315</u>	<u>+1,0,315</u>
n_f	0	0	0,315	0,315	0,315

$$m_o = 0,315 \text{ mol}$$

$$\downarrow \times M = 40 + 12 + 3 \cdot 16$$

$$= 400 \text{ g/mol}$$

$$m = 31,5 \text{ g}$$



$$m = 2,64 \text{ g}$$

$$\downarrow M = 12 + 2 \cdot 16$$

$$= 44 \text{ g/mol}$$

$$m = 0,05$$



$$m = 2,64 \text{ g}$$

$$\downarrow / M = 44 \text{ g/mol}$$

$$n_p = 0,06 \text{ mol}$$

n_0	0,01	0,06	0	0
Δn	+1,0,01	+6,0,01	+6,0,01	+6,0,01
n_f	0	0	0,06	0,06

$$n_0 = 0,01 \text{ mol}$$

$$\downarrow \times M = 6 \cdot 12 + 12 + 6 \cdot 16$$

$$= 180 \text{ g/mol}$$

$$m = 1,8 \text{ g}$$

$$n_0 = 0,06 \text{ mol}$$

$$\downarrow \times M = 32 \text{ g/mol}$$

$$m = 1,92 \text{ g}$$

$$n_p = 0,06 \text{ mol}$$

$$\downarrow \times M = 2 + 16 = 18 \text{ g/mol}$$

$$m = 1,08 \text{ g}$$

masse totale de réactifs
3,72g

masse totale des produits
3,72g

matière conservée
selon Lavoisier rien ne se perd rien ne se crée.